



Chemistry Review: Name _____

17.1-17.6 (pgs 289-303)

18.4 (pgs 317-318)

20.1 & 20.4 (pgs 339-341 & 346)

21.1-21.6 (pgs 352-367)

23.1-23.4 (pgs 388-400)

25.1-25.3 (pgs 425-436)

1. What is an element? 1. _____

2. What is the lightest, most abundant element? 2. _____
3. Write the names of the first 8 elements. 3. _____

4. Write the *symbols* for the first 8 elements. 4. _____

5. Who is Demitri Mendeleev? 5. _____

6. How are the elements grouped on Mendeleev's Periodic Table? 6. _____

7. What does periodic mean? 7. _____
8. What is an atom? 8. _____

9. Why are models of atoms useful? 9. _____

10. What is the Bohr model of an atom? 10. _____

11. What is the present model of the atom? 11. _____

12. Why is an atom electrically neutral? 12. _____

13. Name the 3 major particles in an atom. 13. _____

14. What is the atomic number of an element? 14. _____

15. How are the elements in the periodic table arranged? 15. _____

16. What is an electron cloud and what are energy levels in an atom? 16. _____

17. What is the mass of a proton? 17. _____
18. How is the number of neutrons determined? 18. _____

19. What is the meaning of U-235? 19. _____

20. What does $\begin{matrix} 35 \\ 17 \end{matrix} \text{Cl}$ mean? 20. _____

21. What is a *group* of elements? 21. _____

22. What group of elements has 7 electrons in their outermost energy level? 22. _____
23. How are lithium, sodium, and potassium similar in their electron configuration? 23. _____
24. What are the elements located in columns 3-12 on the periodic table called? 24. _____
25. What is a period on the periodic table and what does the period tell about the elements? 25. _____

26. How many valence electrons to metals have? 26. _____
27. List 3 properties of a metal. 27. _____
28. How many valence electrons do nonmetals have? 28. _____
29. What is a metalloid/semi-conductor? 29. _____
30. What is the most stable arrangement of electrons in groups of atoms? 30. _____
