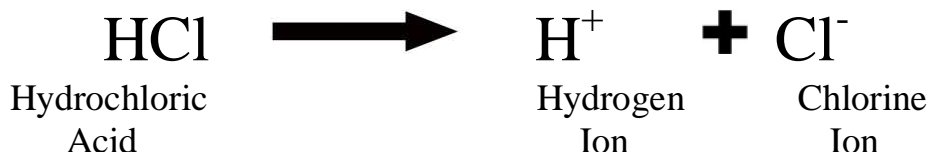




# Acids and Bases

Acids Donate Protons, Bases Accept Them

1. **Acids** (from Latin *acidus* which means “sour”)  
compounds that release hydrogen ions ( $H^+$ ) into a solution



Examples:  $H_2SO_4$  is sulfuric acid     $CH_3COOH$  acetic (vinegar)  
 $HNO_3$  is nitric acid     $H_2CO_3$  is carbonic acid

2. **Bases** (alkaline from Arabic *al-qali* which means “ashes”)  
compounds that release hydroxide ions ( $OH^-$ ) into a solution



Examples:  $KOH$  is potassium hydroxide  
 $Ca(OH)_2$  is calcium hydroxide

3. **Neutralization Reactions** occur when an acid chemically reacts with a base forming ionic salts and often water.



4. **pH Scale** a numeric scale used to express the acidity of a solution

