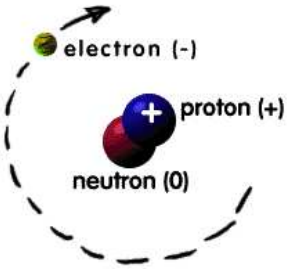
Finding the Parts of an Atom



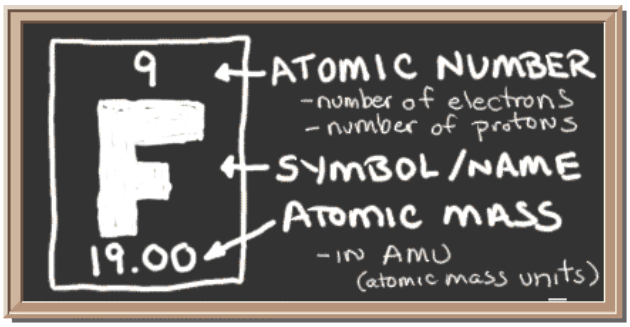
Particle	Charge	Location	Mass
Proton	(+)	Nucleus	1 A.M.U.
Neutron	No Charge	Nucleus	1 A.M.U.
Electron	(-)	Cloud	1/1800 A.M.U. “

Atomic **M**ass **U**nit = .00000000000000000000000000016726 g

Question: If all atoms are made Protons, Neutrons and Electrons…what makes each atom different?

Answer: The Atomic Number

Atomic #	Element	# of Protons
1		
		6
	Fluorine	



1. **Protons**

   The number of protons is equal to the Atomic #

   “Atomic # is 9… so Fluorine has 9 (+)protons”

2. **Electrons**

   Atoms have no net charge so (+)protons = (-)electrons

   “Fluorine has 9 protons… so it has 9 electrons“

3. **Neutrons**

   Mass Number is the mass of the nucleus (protons and neutrons)

   “Fluorine’s mass is 19…so 19 – 9 protons = 10 neutrons”